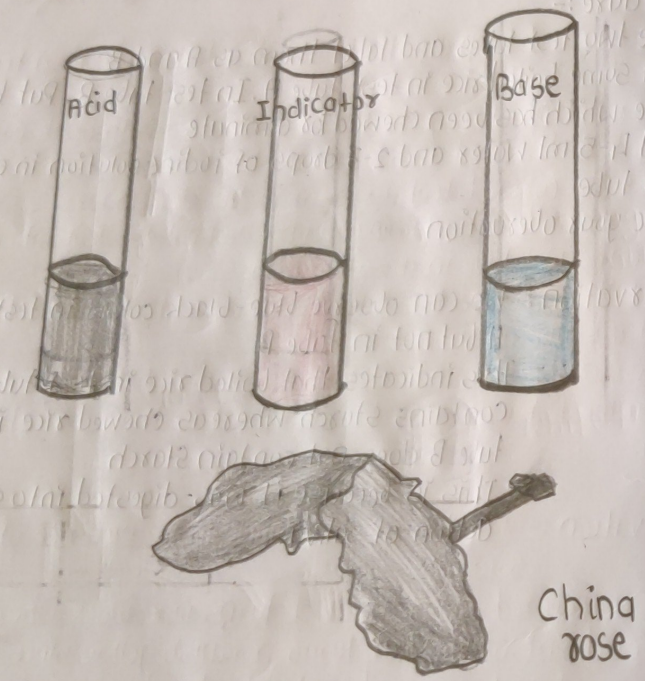


EXPT. NO. 3 NAME:- China rose flower and indicator prepared

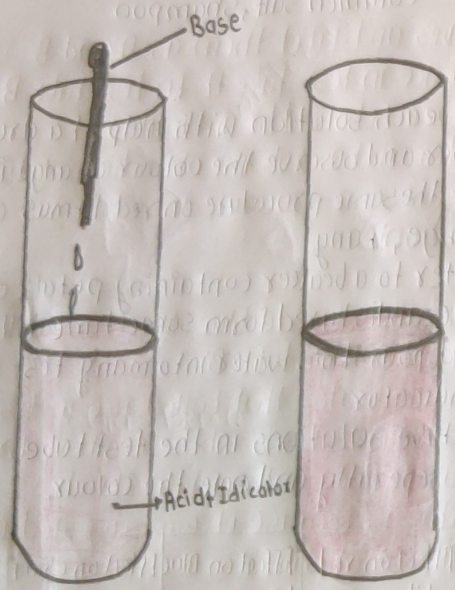
Page No.
Date:
YOUVA

- Aim:- To Find out the chemical nature of substance using litmus paper and china rose solution
- Material required:- Dropper, test tubes, chinarose, blue and red litmus paper, lemon juice, washing soda, commen salt, shampoo
- procedure:-

① Place a drop of each solution with help of a dropper on Blue litmus paper and observe the colour change if any
② Similarly, repeat the same procedure on red litmus and observe the colour change, if any
③ Add warm water to a beaker containing petals of china rose. keep the mixture undistured form some time till the water becomes coloured pour this water into many test tubes. use this water as an indicator
④ Add the test ~~tube~~ solutions in the test tubes containing the indicator seperatly and note the colour.

Test solution	Effect on red Litmus paper	Effect on Blue litmus paper	Effect on china Rose solution	Solution: Acidic/ Basic

Teacher's Signature:



EXPT. NO. 4 NAME: Neutralisation reaction

Page No. Date: YOUVA

- Aim:- To demonstrate neutralisation reation
- Material required :- Dilute hydochoric acid, phenolpthalein, dilute sodium hydroxide, dropper and test tube
- Procedure :- ① take 5ml of dilute hydrochloric acid in a test tube
  ② Add 2-3 drops of phenolpthalin indicator to it
  ③ Shake the test tube. The solution remains colouredless
  ④ Now add dilute sodium hydroxide dropwise till pink Colour just appears
  ⑤ Touch the test tube immediately and check whether it is warm or not
  ⑥ Again add a drop of dilute ~~sodium~~ hydro~~xide~~ chloric acid. the colour pink dissappears
  ⑦ Now add a drop of dilute ~~solution~~ sodium hydroxide the pink colour will appear again
- Observation :- phenolphthalin indicator is colourless in acid medium and pink in basic medium. When acid and base are mixed, by neutralised each other and and the solution formed is neither acidic nor basic. Adding a single drop of base to the neutrailised solution turns it pink similarly adding a single drop of acid to the same solution turns it colourless

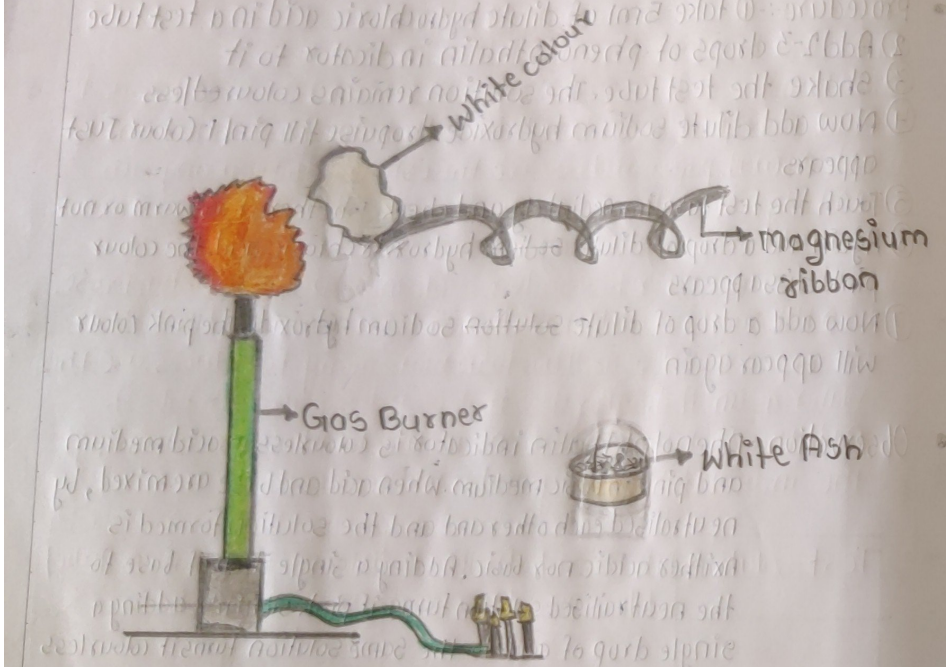
The reaction that takes place is as Follow

$$HCL + NaOH \longrightarrow NaCl + H_2O$$

[Hydrochloric acid] [sodium hidroxide] [Sodium chloride (salt)] [water]

The test tube feels warm when touch because heat is evolved during a nutralisation reaction

Teacher's Signature: _____



EXPT. NO. 5 | NAME | Page No. | Date: | YOUVA

- **Aim:-** To observe a chemical change during burning of magnesium ribbon

- **Note:-** perform this activity only under the supervision of your Teacher

- **Material required:-** Magnesium ribbon, sand paper, a pair of tongs and burner

- **procedure:-**
  - Take a small pieces of a magnesium ribbon and clean its tip with sandpaper
  - Hold it with a pair of tongs over the flame of a burner
  - It burns with a brilliant white light and a white powdery ash of magnesium oxide (MgO) is formed
  - It is a new substance and is different from magnesium ribbon
  - Take a small amount of this ash and dissolve it in water in a test tube. Stir it properly
  - Test this solution with a blue litmus paper and a red litmus paper

- **Observation:-**
  On burning the magnesium ribbon, a new substance MgO is formed.
  On mixing MgO with water, it forms another new substance $Mg(OH)_2$. This substance is basic in nature as it turns red litmus paper blue. Therefore, burning of magnesium ribbon is an example of a chemical change. The reactions that take place are as following.

Teacher's Signature: _____

EXPT. NO. | NAME | Page No. | Date | YOUVA

Magnesium + Oxygen → magnesium oxide [ash]
(2Mg) (O₂) (2MgO)

Magnesium oxide + Water → magnesium hydroxider
(MgO) H_2O $Mg(OH)_2$

caution:- Do not stare at the burning ribbon for long. it is harmful to the eyes. Use glasses to look at it.

Teacher's Signature: